Electronic configuration of chromium is

Options:

[Ar]3d⁵4s¹
[Ar]3d⁴4s²
[Ar]3d³4s²
[Ar]3d⁵4s²
In their compounds, the maximum oxidation states attained by Fe and Os respectively are
Options:
+VIII and + VI
+VI and + VIII
+II and + VI
+III and + VIII

Question Number : 3  Question Id : 827347703  Question Type : MCQ

The spin only magnetic moment of CuSO$_4$.5H$_2$O in Bohr magneton (BM) unit is
Options:
2.83
3.87
1.73
4.90

Question Number : 4  Question Id : 827347704  Question Type : MCQ

Bright orange color of Chromium trioxide (or Chromic acid) arises from
Options:
Presence of impurities
f-f transitions
d-d spectra
Charge transfer

Question Number : 5  Question Id : 827347705  Question Type : MCQ

$[\text{Mn EDTA}]^{4-}$ Complex has very pale color because d-d transition is forbidden by the Laporte selection rule. This states that when promoting an electron the change in the subsidiary quantum number (l) must be
Options:
+1
-1
±1
0

Question Number : 6  Question Id : 827347706  Question Type : MCQ

Electronic configuration of gadolinium (Gd) is
Options:
[Xe]4f$^7$5d$^1$6s$^2$
[Xe]4f$^7$6s$^2$
[Xe]4f$^8$6s$^2$
[Xe]4f$^9$6s$^2$
Even though Ln$^{3+}$ dominates the chemistry of lanthanide elements, still Eu$^{2+}$ exists because oxidation number (+II) lead to

Options:
Noble gas configuration
Half filled $f^*$ shell
Completely filled $f^*$ level
Completely filled $d^*$ level

Question Number : 8  Question Id : 827347708  Question Type : MCQ

Yellow colors of Sm$^{3+}$ and Dy$^{3+}$ is because number of unpaired electrons in both are

Options:
3
4
5
6

Question Number : 9  Question Id : 827347709  Question Type : MCQ

Separation of Dy$^{3+}$ and Ho$^{3+}$ is very difficult because they have similar

Options:
Physical state
Color
Charge
sizes

Question Number : 10  Question Id : 827347710  Question Type : MCQ

Cyanides, nitrites and carbonyl compounds can be reduced by using

Options:
$B_2H_5$
$B_4H_{10}$
$B_6H_{10}$
$B_9H_{15}$

Question Number : 11  Question Id : 827347711  Question Type : MCQ

Borazine is sometimes called inorganic benzene. However, in comparison to benzene, borazine’s reactivity is

Options:
Less
More
Equal
Very less
Question Number : 12  Question Id : 827347712  Question Type : MCQ

Total number of three-centre two-electron (3c-2e) ‘banana-shaped’ bonds in diborane are
Options:
Zero
One
Two
Three

Question Number : 13  Question Id : 827347713  Question Type : MCQ

Carborane is formed by reaction of decaborane-14 with
Options:
Methane
Ethane
Ethene
Ethyne

Question Number : 14  Question Id : 827347714  Question Type : MCQ

Cement, ceramic and glass industries are based on the chemistry of
Options:
Silicates
Diborane
Borazine
Silicones

Question Number : 15  Question Id : 827347715  Question Type : MCQ

To detect end point in acid-base neutralization reactions occurring in aqueous solvent (H₂O) and non-aqueous solvents (liquid ammonia), suitable indicator is
Options:
Calmagite
Phenolphthalein
Erio-chrome black T
KMnO₄

Question Number : 16  Question Id : 827347716  Question Type : MCQ

In N₂O₄ as solvent, NOCl is an acid since it produces
Options:
H⁺
H₃O⁺
NO⁺
NO₃⁻
Question Number : 17  Question Id : 827347717  Question Type : MCQ

Reaction of TiBr₄ and N₂O₄ give following nitrato complex
Options :
  Ti(NO₃)₄
  Ti(NO₃)₆
  Ti(NO₃)₂
  Ti(NO₃)₅

Question Number : 18  Question Id : 827347718  Question Type : MCQ

Mineral acids (HNO₃, H₂SO₄ and HCl) in HF (non-aqueous solvent) behave as
Options :
  Solvents
  Solutes
  Acids
  bases

Question Number : 19  Question Id : 827347719  Question Type : MCQ

Using AlEt₃/TiCl₄ catalyst, polymerization of propene produces polymer which is
Options :
  Stereo regular
  Atactic
  Random
  Cross-linked

Question Number : 20  Question Id : 827347720  Question Type : MCQ

In vanadocene, the number of carbon atoms in each ring bonded to vanadium are
Options :
  4
  5
  3
  10

Question Number : 21  Question Id : 827347721  Question Type : MCQ

Reaction of ferrocene with an equimolar amount of acetyl chloride gives a product in which cyclopentadienyl rings undergo
Options :
  Diels-Alder reaction
  Catalytic hydrogenation
  Friedel-Crafts acylation
  Hydroformylation
Question Number : 22  Question Id : 827347722  Question Type : MCQ

Reaction of 1 mole of organocadmium compound with 2 moles of acetyl chloride gives 2 moles of
Options :
Alcohols
Carboxylic acids
Aldehydes
Ketones

Question Number : 23  Question Id : 827347723  Question Type : MCQ

As per the liquid drop model for the nuclear model, the range of attractive forces is
Options :
1 fin to 2 fin
2 fin to 3 fin
3 fin to 4 fin
4 fin to 5 fin

Question Number : 24  Question Id : 827347724  Question Type : MCQ

The nuclei of C atoms which have had the orbital electrons removed are called
Options :
Carbene
Carbocation
Stripped carbon
Carboanion

Question Number : 25  Question Id : 827347725  Question Type : MCQ

The total mass of the fission products is some 0.22 mass units less than the mass of the uranium atom and neutron. This corresponds to an energy release of over
Options :
100 MeV
200 MeV
300 MeV
400 MeV

Question Number : 26  Question Id : 827347726  Question Type : MCQ

Mass of proton is
Options :
1.008 665 amu
1.007 825 amu
1.007 277 amu
1.009 235 amu

Question Number : 27  Question Id : 827347727  Question Type : MCQ
Which one of the following is not the application of radioisotopes
Options:
Estimation of the age of various objects
Estimation of the age of mineral deposits
Determination of solubility of sparingly soluble materials
Preparation of detergents

Question Number : 28  Question Id : 827347728  Question Type : MCQ

Presence of which one of the following salt does not contribute hardness in water
Options:
Sodium bicarbonate
Calcium bicarbonate
Magnesium sulphate
Magnesium chloride

Question Number : 29  Question Id : 827347729  Question Type : MCQ

Which one of the following is not used to soften water
Options:
Ion-exchange methods
Addition of Grahams salt (Calgon)
Addition of sodium carbonate
Addition of sugar

Question Number : 30  Question Id : 827347730  Question Type : MCQ

In water, harmful concentration of fluoride ions is
Options:
Above 2 ppm
Below 1 ppm
Below 0.5 ppm
Below 0.7 ppm

Question Number : 31  Question Id : 827347731  Question Type : MCQ

Completely biodegradable detergent is
Options:
Linear alkylbenzene sulphonates
Teepol
Branched chain alkylbenzene sulphonates
Calcium stearate

Question Number : 32  Question Id : 827347732  Question Type : MCQ
Which one of the following does not occur by the use of biodegradable detergents
Options:
Decrease in concentration of dissolved dioxygen
Deaths of fish or plants
Unpleasant smells
Formation of insoluble precipitate or ‘scum’ when water is hard

Question Number : 33  Question Id : 827347733  Question Type : MCQ

‘Soda glass’ is made by fusing the SiO₂, CaCO₃ and
Options:
Na₂CO₃
K₂CO₃
CaO
PbO

Question Number : 34  Question Id : 827347734  Question Type : MCQ

Borosilicate glasses are less prone to chemical attack because they contain less
Options:
CaCO₃
Alkali
SiO₂
B₂O₃

Question Number : 35  Question Id : 827347735  Question Type : MCQ

Alloys of Nb₃Sn, Nb₃Ge, Nb₃Al and V₃Si all show superconductivity and have critical temperature (Tc) values of about
Options:
10 K
20 K
30 K
40 K

Question Number : 36  Question Id : 827347736  Question Type : MCQ

The compound YBa₂Cu₃O₇₋ₓ was the first superconductor which worked at 93K. The temperature was important because it allows liquid nitrogen to be used as coolant rather than the more expensive liquid helium. This Y-Ba-Cu-O system is called
Options:
1-2-3 system
2-3-4 system
3-4-5 system
4-5-6 system
Question Number : 37  Question Id : 827347737  Question Type : MCQ

The IUPAC name of tert-Butyl alcohol is
Options :
2-methyl-2-propanol
2-methyl-2-propanal
2-methyl-1-propanol
2-methyl-1-propanal

Question Number : 38  Question Id : 827347738  Question Type : MCQ

That part of the science which deals with the molecular structure in three dimensions is called
Options :
Organic chemistry
Stereo chemistry
Inorganic chemistry
Physical chemistry

Question Number : 39  Question Id : 827347739  Question Type : MCQ

To specify a particular configuration in some simpler, easier way than by always having to draw its picture, prefix R or S is used. As per sequence rules, the complete sequence of priority for sec-butyl chloride is
Options :
H, CH₃, C₂H₅, Cl
CH₃, H, Cl, C₂H₅
Cl, C₂H₅, CH₃, H
Cl, H, CH₃, C₂H₅

Question Number : 40  Question Id : 827347740  Question Type : MCQ

The maximum number of stereoisomers that can exist for compounds like glucose which contain five chiral centers is
Options :
Four
Eight
Sixteen
Thirty two

Question Number : 41  Question Id : 827347741  Question Type : MCQ

Resolution of a racemic modification means the separation of a racemic modification into
Options :
Enantiomers
Diastereomers
Conformers
Stereoisomers

Question Number : 42  Question Id : 827347742  Question Type : MCQ

Photolysis of diazomethane in liquid cis-2-butene gives only cis-1,2-dimethylcyclopropane, and in liquid trans-2-butene gives only trans-1,2-dimethylcyclopropane. Cycloaddition in this example is not

Options :
Stereoselective
Stereospecific
Syn
Anti

Question Number : 43  Question Id : 827347743  Question Type : MCQ

The simultaneous making of two bonds on the opposite faces of a component is known as

Options :
Suprafacial addition
Antarafacial addition
Cycloreversion
Cyclosubtraction

Question Number : 44  Question Id : 827347744  Question Type : MCQ

The [R]-rearrangement of hexa-1,5-diene is sigmatropic reaction of the type designated as

Options :
[1,3]
[1,6]
[3,3]
[6,6]

Question Number : 45  Question Id : 827347745  Question Type : MCQ

In electrocyclic reaction, cis-3,4-dimethylcyclobutene gives

Options :
cis,trans-2,4-hexadiene
trans,trans-2,4-hexadiene
cis,cis-2,4-hexadiene
2,4-hexadiene

Question Number : 46  Question Id : 827347746  Question Type : MCQ
Conrotatory motion in thermal cyclization of cis.trans-2,4-hexadiene gives
Options:
3,4-Dimethylcyclobutene
cis-3,4-Dimethylcyclobutene
trans-3,4-Dimethylcyclobutene
1,2-Dimethylcyclobutene

Question Number : 47  Question Id : 827347747  Question Type : MCQ

2-methyl-2-butene on hydroboration-oxidation gives
Options:
3-methyl-2-butanol
2-methyl-2-butanol
2-methyl-butene
3,3-dimethyl-1-butanol

Question Number : 48  Question Id : 827347748  Question Type : MCQ

n-propyl alcohol can be converted to propionaldehyde by using
Options:
Pyridinium chlorochromate
Potassium permanganate
Potassium dichromate
CrO₃

Question Number : 49  Question Id : 827347749  Question Type : MCQ

Hydrolysis of m-chlorobenzene diazonium hydrogen sulfate gives
Options:
α-Chlorophenol
m-Chlorophenol
p-Chlorophenol
phenol

Question Number : 50  Question Id : 827347750  Question Type : MCQ

Reaction of phenol with propionyl chloride gives phenyl propionate. The product on heating with aluminium chloride undergoes Fries rearrangement to produce
Options:
α-Hydroxyphenyl ethyl ketone
p-Hydroxyphenyl ethyl ketone
Mixture of α- and p- Hydroxyphenyl ethyl ketone
m-Hydroxyphenyl ethyl ketone
Question Number : 51  Question Id : 827347751  Question Type : MCQ

Clemmensen reduction of n-pentyl phenyl ketone with amalgamated zinc and conc. HCl gives
Options :
Hexane
Benzene
n-hexylbenzene
n-pentyltoluene

Question Number : 52  Question Id : 827347752  Question Type : MCQ

Wolff-Kishner reduction of n-propyl m-tolyl ketone using hydrazine and base gives
Options :
Butane
Toluene
Benzene
m-(n-butyl) toluene

Question Number : 53  Question Id : 827347753  Question Type : MCQ

Lithium dimethylcuprate react easily with p-nitrobenzoyl chloride to give
Options :
Methyl phenyl ketone
Methyl p-nitrophenyl ketone
Acetophenone
p-nitrobenzene

Question Number : 54  Question Id : 827347754  Question Type : MCQ

In the presence of Conc. NaOH, a mixture of p-methoxybenzaldehyde and formaldehyde gives
Options :
p-methoxybenzyl alcohol
Sodium formate
Mixture of p-methoxybenzyl alcohol and Sodium formate
methanol

Question Number : 55  Question Id : 827347755  Question Type : MCQ

Reaction of phenylmagnesium bromide with ethylene oxide gives a product which on addition of water gets converted into
Options :
2-phenylethanol
Phenol
Ethanol
2-cyclohexylethanol

**Question Number : 56  Question Id : 827347756  Question Type : MCQ**

The strongest acid amongst acetic acid, chloroacetic acid, dichloroacetic acid, and trichloroacetic acid is
Options:
- acetic acid,
- chloroacetic acid
- dichloroacetic acid
- trichloroacetic acid

**Question Number : 57  Question Id : 827347757  Question Type : MCQ**

Oxidation of p-nitrotoluene with hot acidic K$_2$Cr$_2$O$_7$ gives
Options:
- Benzaldehyde
- Benzoic acid
- p-nitrobenzoic acid
- p-nitrobenzaldehyde

**Question Number : 58  Question Id : 827347758  Question Type : MCQ**

Hydrolysis of o-Tolunitrile using 75% H$_2$SO$_4$ at 150-160 °C gives
Options:
- o-Toluene
- o-Toluic acid
- Benzene
- Benzoic acid

**Question Number : 59  Question Id : 827347759  Question Type : MCQ**

Reduction of m-Toluic acid using LiAlH$_4$ gives
Options:
- m-methylbenzyl alcohol
- Benzyl alcohol
- Toluene
- m-methyl phenol

**Question Number : 60  Question Id : 827347760  Question Type : MCQ**

Ethyl acetoacetate can be converted to 5-methyl-2-hexanone by acetoacetic ester synthesis of ketone using
Options:
- Butyl bromide
Isobutyl bromide
Propyl bromide
Ethyl bromide

Question Number : 61  Question Id : 827347761  Question Type : MCQ

The increasing order of basic strength amongst following amines (methylamine, dimethylaniline, trimethylaniline and aniline) is
Options :
Aniline < Trimethylamine < methylamine < Dimethylamine
Aniline < methylamine < Dimethylamine < Trimethylamine
Aniline < Trimethylamine < Dimethylamine < methylamine
methylamine< Dimethylamine < Trimethylamine < Aniline

Question Number : 62  Question Id : 827347762  Question Type : MCQ

Reduction of p-nitroaniline using Sn, HCl and heat gives
Options :
p-nitrobenzene
p-phenylenediamine
Aniline
benzene

Question Number : 63  Question Id : 827347763  Question Type : MCQ

Reductive amination of acetone with NH₃, H₂ and Ni gives
Options :
Methylamine
Propylamine
Isopropylamine
Ethylamine

Question Number : 64  Question Id : 827347764  Question Type : MCQ

Hofmann degradation of m-bromobenzamide using KOBr gives
Options :
Aniline
Bromobenzene
Benzene
m-bromoaniline

Question Number : 65  Question Id : 827347765  Question Type : MCQ

To produce the molecular ion (or parent ion) in a mass spectrum, the number of electrons removed from the parent molecule are
Options :
Question Number : 66  Question Id : 827347766  Question Type : MCQ

In the infrared spectrum, a strong, broad band in the 3200-3600 cm\(^{-1}\) region \{due to O-H stretching\} and another strong, broad band, \{due to C-O stretching\} were observed in the 1000-1200 cm\(^{-1}\) region. The organic molecule is

Options:
- Phenol
- Hydrogen-bonded alcohol
- Ether
- Carboxylic acid

Question Number : 67  Question Id : 827347767  Question Type : MCQ

The correct increasing order w.r.t. energy required for \(\pi\rightarrow\pi^*\) transition in the ultraviolet spectrum amongst ethylene, 1,3-butadiene and \(\beta\)-carotene is

Options:
- Ethylene \(<\) 1,3-butadiene \(<\) \(\beta\)-carotene
- 1,3-butadiene \(<\) ethylene \(<\) \(\beta\)-carotene
- \(\beta\)-carotene \(<\) 1,3-butadiene \(<\) ethylene
- \(\beta\)-carotene \(<\) ethylene \(<\) 1,3-butadiene

Question Number : 68  Question Id : 827347768  Question Type : MCQ

One should look at which aspect of the NMR spectrum to find out how many protons of each kind are there in a molecule

Options:
- The number of signals
- The positions of signals
- The splitting of a signal
- The intensities of the signals

Question Number : 69  Question Id : 827347769  Question Type : MCQ

Photochemistry is the study of chemical reactions resulting from the exposure of

Options:
- Light radiations
- Swift heavy ions
- Nuclear radiations
- Ultrasonic waves

Question Number : 70  Question Id : 827347770  Question Type : MCQ
Upon absorption of radiation, changes occurring in molecular energy levels is schematically represented by

Options:
Norrish diagram
Jablonski diagram
Barton diagram
Paterno-Buchi diagram

Question Number : 71  Question Id : 827347771  Question Type : MCQ

Non-radiative transition so that energy of activated molecule may be lost is

Options:
Fluorescence
Phosphorescence
Internal conversion
Chemiluminescence

Question Number : 72  Question Id : 827347772  Question Type : MCQ

Wavelengths of the emitted radiations in fluorescence compared to the wavelength of exciting radiations are

Options:
Equal
Much shorter
Shorter
greater

Question Number : 73  Question Id : 827347773  Question Type : MCQ

Analysis of the X-ray diffraction pattern makes it possible to measure the interatomic distance between any two nearby atoms in a crystal. When a beam of light is scattered by an object containing regularly spaced atoms in a crystal, diffraction of electromagnetic radiation occurs. This scattering can happen only if the wavelength of the radiation is “A” the spacing between the atoms.

What is “A” in the above statement?

Options:
Comparable to
Greater than
Less than
Not equal to

Question Number : 74  Question Id : 827347774  Question Type : MCQ

Unit cell of NaCl and KCl is of the following type

Options:
bcc
fcc
hcc
Primitive-cubic

Question Number : 75  Question Id : 827347775  Question Type : MCQ

Graphite is a covalent network solid consisting of two-dimensional sheets of carbon atoms organized into six-membered rings. The hybridization of carbon atoms in graphite is

Options :
- \( sp \)
- \( sp^3 \)
- \( sp^2 \)
- \( sp^3d \)

Question Number : 76  Question Id : 827347776  Question Type : MCQ

It is the hardest known substance. In addition to its use in jewelry, it is widely used industrially for the tips of saw blades and drilling bits. It is an electrical insulator. It is one of the carbon allotropes. What is the name of this allotrope of carbon?

Options :
- \( C_{60} \)
- Fullerene
- Graphite
- Diamond

Question Number : 77  Question Id : 827347777  Question Type : MCQ

When hot water is discharged from industrial plants, thermal pollution in lakes and rivers take place. The important consequence is the damage to aquatic life that can result from this thermal pollution. What happens to the concentration of dissolved oxygen \([DO]\), when hot water is discharged in lakes and rivers?

Options :
- \([DO]\) decreases
- \([DO]\) increases
- \([DO]\) remains unaffected
- \([DO]\) becomes zero

Question Number : 78  Question Id : 827347778  Question Type : MCQ

At a given temperature, the solubility of a gas in a liquid is directly proportional to the partial pressure of the gas over the solution. This is known as

Options :
- Raoult’s law
Henry's law
Nernst distribution law
Joule's law

Question Number : 79  Question Id : 827347779  Question Type : MCQ
Which of the following practical uses is not based on colligative properties
Options :
- Melting of snow by salt
- The desalination of sea water by reverse osmosis
- Continuous bubbling of carbonated drinks as they warm up to room temperature after being refrigerated
- The separation & purification of volatile liquids by fractional distillation.

Question Number : 80  Question Id : 827347780  Question Type : MCQ
If a solution contains a non-volatile solute, its vapor pressure is equal to the product of vapor pressure of the pure solvent with the mole fraction of the solvent. This law is known as
Options :
- Raoult's law
- Henry's law
- Nernst distribution law
- Joules law

Question Number : 81  Question Id : 827347781  Question Type : MCQ
The sign of entropy change for the process of sublimation of dry ice is
Options :
- Positive
- Negative
- Zero
- Uncertain

Question Number : 82  Question Id : 827347782  Question Type : MCQ
According to third law of thermodynamics, at zero Kelvin, the entropy of a crystalline, perfectly ordered substance is
Options :
- 4
- 2
- 1
- 0

Question Number : 83  Question Id : 827347783  Question Type : MCQ
“Energy can never be created nor be destroyed, but it can be changed from one form to another”. This statement is known as
Options:
First law of thermodynamics
Second law of thermodynamics
Third law of thermodynamics
Fourth law of thermodynamics

Question Number : 84  Question Id : 827347784  Question Type : MCQ

Entropy change is negative for which one of the following processes:
Options:
Phase transitions that convert a solid to a liquid
Reactions that increase the number of gaseous molecules
Lowering the temperature of a substance
Dissolution of molecular solids in water

Question Number : 85  Question Id : 827347785  Question Type : MCQ

A wave function contains all the dynamical information about a system and is found by solving the appropriate
Options:
Schrodinger equation
Debye-Huckel-Onsager equation
Bragg’s equation
Arrhenius equation

Question Number : 86  Question Id : 827347786  Question Type : MCQ

The energy levels of a particle of mass ‘m’ in a box of length ‘L’ are
Options:
Non-quantized
Quantized
Similar
Dissimilar

Question Number : 87  Question Id : 827347787  Question Type : MCQ

When a particle is subjected to a restoring force proportional to the displacement, it undergoes harmonic motion. This is known as
Options:
Bragg’s law
Nernst law
Hooke’s law
Arrhenius’s law
The energy levels of a harmonic oscillator are equally spaced and specified by the quantum number whose values are
Options:
0, 5, 10, ....
2, 4, 6, ....
1, 3, 5, ....
0, 1, 2, ....

Question Number: 89  Question Id: 827347789  Question Type: MCQ

Hydrogen electrode is the type of electrode best described as a
Options:
Gas electrode
Metal-insoluble-salt electrode
Redox electrode
Glass electrode

Question Number: 90  Question Id: 827347790  Question Type: MCQ

What range should a voltmeter have (in volts) to display changes of pH from 1 to 14 at 25 °C if it is arranged to give a reading of zero when pH = 7?
Options:
A range of 0.66 V
A range of 0.77 V
A range of 0.88 V
A range of 0.99 V

Question Number: 91  Question Id: 827347791  Question Type: MCQ

According to Nernst equation, if the reaction quotient for the cell reaction is decreased, then the tendency of a reaction to form products becomes
Options:
Zero
10 times
Greater
Smaller

Question Number: 92  Question Id: 827347792  Question Type: MCQ

For calcium hydroxide, solubility (S) and Solubility product (K) are related as
Options:
\[ K = 4S^3 \]
\[ K = 4S \]
\[ K = S^2 \]
\[ K = 2S^3 \]

Question Number : 93  Question Id : 827347793  Question Type : MCQ

For a unimolecular reaction, ten times as many reactant molecules decay in a given time interval when there are initially 50 reactant molecules as when there are only 5 reactant molecules present. This unimolecular reaction is
Options :
Zero order
First order
Second order
Third order

Question Number : 94  Question Id : 827347794  Question Type : MCQ

According to Michaelis Menten kinetics for enzyme catalysed reaction, the rate of formation of product is maximum when concentration of substrate is
Options :
Low
High
Unity
fractional

Question Number : 95  Question Id : 827347795  Question Type : MCQ

Suppose ‘c’ is the intercept and ‘m’ is the slope of a straight line when a graph of the reciprocal of the reaction rate is plotted against the reciprocal of the substrate concentration. The value of Michaelis constant can be calculated by using the equation
Options :
\( \frac{m}{c} \)
\( \frac{c}{m} \)
mc
mc
m

Question Number : 96  Question Id : 827347796  Question Type : MCQ

Reactions in which an intermediate formed in one step generates a reactive intermediate in the next step, then that intermediate generates another reactive intermediate, and so on, are known as
Options :
Enzyme catalyzed reactions
First order reactions
Chain reactions
Unimolecular reactions

Question Number : 97  Question Id : 827347797  Question Type : MCQ

The source of chlorofluoro carbons is
Vehicle’s exhaust
Fossil fuel burning
Deforestation
Refrigerants

Question Number : 98  Question Id : 827347798  Question Type : MCQ

In water, presence of energy or any foreign substance in such concentration and for such duration that tends to degrade the quality of water so that users (humans, animals or any other organism) cannot enjoy the beneficial qualities of water but the use constitutes a hazard is known as

Options :
Air pollution
Water pollution
Thermal pollution
Noise pollution

Question Number : 99  Question Id : 827347799  Question Type : MCQ

Which one of the following is not a greenhouse gas

Options :
CO₂
CH₄
N₂
CFC

Question Number : 100  Question Id : 827347800  Question Type : MCQ

pH value of acid rain is

Options :
10.2-10.4
8.2-8.3
7.2-7.6
4.1-4.3